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## Molar Mass Answer Key

**molar mass worksheet answer key - riverdell** - molar mass worksheet - answer key calculate the molar masses of the following chemicals: 1)  $\text{Cl}_2$  2)  $\text{KOH}$  3)  $\text{BeCl}_2$  4)  $\text{FeCl}_3$  5)  $\text{BF}_3$  6)  $\text{CCl}_2\text{F}_2$  7)  $\text{Mg}(\text{OH})_2$  8)  $\text{UF}_6$  9)  $\text{SO}_2$  10)  $\text{H}_3\text{PO}_4$  11)  $(\text{NH}_4)_2\text{SO}_4$  12)  $\text{CH}_3\text{COOH}$  13)  $\text{Pb}(\text{NO}_3)_2$  14)  $\text{Ga}_2(\text{SO}_3)_3$

**molar mass worksheet - nclark** - molar mass worksheet calculate the molar mass of the following chemicals: 1)  $\text{Cl}_2$  2)  $\text{KOH}$  3)  $\text{BeCl}_2$  4)  $\text{FeCl}_3$  5)  $\text{BF}_3$  6)  $\text{CCl}_2\text{F}_2$  7)  $\text{Mg}(\text{OH})_2$  8)  $\text{UF}_6$  9)  $\text{SO}_2$  10)  $\text{H}_3\text{PO}_4$  11)  $(\text{NH}_4)_2\text{SO}_4$  12)  $\text{CH}_3\text{COOH}$  13)  $\text{Pb}(\text{NO}_3)_2$  14)  $\text{Ga}_2(\text{SO}_3)_3$

**molar mass practice worksheet - my chemistry class** - solutions to the molar mass practice worksheet: important note to students: all of the units given here are "grams per mole", which may be abbreviated as "g/mol", "grams/mol", or "g . mol<sup>-1</sup>", depending on how your teacher likes to see it written. they all mean the same thing, but it's **relative mass and the mole answer key** - relative mass and the mole answer key created date: 20171005134609z ... **chapter 3 molar mass calculation of molar masses** - molar mass molar ratio  $180.1 \text{ g mol}^{-1} \times 2.33 \text{ mol} = 418.6 \text{ g}$  strategy: once the number of moles of a substance present is known, we can use: - molar mass to find the number of grams - avogadro's number to find the number of atoms, ions, or molecules  $\text{moles} \times \text{grams} = \text{atoms} \times \text{molar mass} = \text{na} \times \text{moles}$  a molar ratio molar conversions chapter 3 h? grams ... **chemistry computing formula mass worksheet - isd 622** - molar ratios the molar ratio is an important concept in solving stoichiometry problems. the sources for these ratios are the coefficients of a balanced equation. example 1:  $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$  what is the molar ratio between  $\text{H}_2$  and  $\text{O}_2$ ? answer: two to one. so this ratio is written as a fraction is what is the molar ratio between  $\text{O}_2$  and  $\text{H}_2\text{O}$ ? answer ... **molar mass name chem worksheet 11-2** - what is its molar mass? 7. find the molar mass of acetylene  $\text{C}_2\text{H}_2$ , the gas used in welding torches. 8. vinegar contains acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ . what is the molar mass of this substance? 9. what is the molar mass of magnesium sulfate,  $\text{MgSO}_4$ ? 10. find the molar mass of phosphorus pentachloride,  $\text{PCl}_5$ . 11. bromine is diatomic. find the molar ... **molar mass problems - incomplete key s o h o** - consider the molecule  $\text{C}_4\text{H}_{10}$  as you answer 11 - 19. 11. name the elements present. 12. how many atoms form the molecule? 13. how many of each atom in the molecule? 14. how many hydrogen atoms in one mole of molecules? 15. how many chlorine atoms in six moles of molecules? 16. what is the molar mass of this molecule? 17. name this molecule. 18. **ws molar mass - saddleback college** - molar mass worksheet and key for the following compounds, write the chemical formula and determine the molar mass. write the units! water carbon dioxide sodium chloride calcium hydroxide potassium permanganate ( $\text{KMnO}_4$ ) sodium carbonate barium nitrate hydrogen monochloride sulfuric acid ( $\text{H}_2\text{SO}_4$ ) acetic acid ( $\text{C}_2\text{H}_4\text{O}_2$ ) see next page for key **mole calculation worksheet - sheffieldschools** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp (you do not need to worry about this yet) each definition can be written as a set of two conversion factors. they are: 1 mole = molar mass(g) can be written as  $\frac{1 \text{ mole}}{\text{molar mass (g)}}$  or  $\frac{\text{molar mass (g)}}{1 \text{ mole}}$

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