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# Mole Concept Homework Packet Answer Key

**homework-2 - mole concept (10 pts)** - homework-2 - mole concept (10 pts) isp 207, f'06 answers note: to obtain full credit, you must (1) show your work, (2) include units in your calculations & (3) express your answers in the correct number of significant figures. 1. **practice problems -mole concept - chemistryu - practice problems-mole concept key [not an assignment]** 1. if a student has avogadro's number of  $\text{CO}_2$  molecules, how many molecules of  $\text{CO}_2$  does he/she have?  $6.02 \times 10^{23}$  molecules of  $\text{CO}_2$  2. (a) how many atoms of c are contained in 1 mole of  $\text{CO}_2$ ? 1 mole  $\text{CO}_2 \times 1 \text{ mole c} = 1 \text{ mole } \text{CO}_2 \times 6.02 \times 10^{23} \text{ atoms c} = 6.02 \times 10^{23} \text{ atoms c}$  (b) how many atoms of o are contained in 1 mole of  $\text{CO}_2$ ? **the mole concept - hcc learning web** - the mole concept 45 section 8.1 avogadro's number 2. element average mass element average mass (a) c 12.01 amu (b) s 32.07 amu (c) se 78.96 amu (d) i 126.90 amu 4. element mass element mass (a) c 12.01 g (b) s 32.07 g (c) se 78.96 g (d) i 126.90 g section 8.2 mole calculations i 6. **worksheet: more mole problems name** - title: microsoft word - 7-11a,b more moles problems wkst-key c author: brent white created date: 6/23/2005 10:43:57 pm **lab activity title: mole day submitted by: recommended ...** - out how i was able to decide i had a mole of a substance and then introduce the concept of molar mass. i then assign a reading assignment in the text over the mole concept as homework. note: if you have large classes you might want to consider having a parent or teacher's aide available to assist you on mole day. it certainly makes the day ... **ap chemistry summer homework - images.pcmac** - ap chemistry summer assignment! welcome to ap chemistry! ap chemistry is a college-level course. you will need to be ... the mole concept and conversions, and ... and therefore the pace is fast. learning from the textbook and doing the homework are the students' responsibilities with the labs serving to supplement the learning in the lecture **molar mass worksheet answer key - river dell regional ...** - molar mass worksheet - answer key calculate the molar masses of the following chemicals: 1)  $\text{Cl}_2$  71 g/mol 2)  $\text{KOH}$  56.1 g/mol 3)  $\text{BeCl}_2$  80 g/mol 4)  $\text{FeCl}_3$  162.3 g/mol 5)  $\text{BF}_3$  67.8 g/mol 6)  $\text{C}_2\text{F}_2$  ... "grams/mole" is better, because any macroscopic amount of a substance is better expressed in grams than amu. **moles worksheet - awesome science teacher resources** - moles worksheet 1) define "mole". 2) how many moles are present in 34 grams of  $\text{Cu(OH)}_2$ ? 3) how many moles are present in  $2.45 \times 10^{23}$  molecules of  $\text{CH}_4$ ? 4) how many grams are there in  $3.4 \times 10^{24}$  molecules of  $\text{NH}_3$ ? 5) how much does 4.2 moles of  $\text{Ca(NO}_3)_2$  weigh? 6) what is the molar mass of  $\text{MgO}$ ? **moles, molecules, and grams worksheet and key** - moles, molecules, and grams worksheet and key 1) how many moles are there in 24.0 grams of  $\text{Fe}$ ? 2) how many moles are there in 458 grams of  $\text{Na}_2\text{SO}_4$ ? 3) how many grams are there in  $2.30 \times 10^{24}$  atoms of silver? 4) how many grams are there in 7.40 moles of  $\text{AgNO}_3$ ? see last page for key . **mole calculations problems review psi chemistry name** - njctl chemistry mole calculations mole calculations problems review psi chemistry name \_\_\_\_\_ the mole and avogadro's number classwork 1. how many iron atoms are in 3 moles of iron? 2. how many sulfur atoms are in 0.25 moles of sulfur? 3. **weekly assignments for chemistry 202** - weekly assignments for chemistry 202 week 1: august 28-september 1 . monday, august 28: read through course requirements, explore the website (and lon-capa) and begin watching **mole calculation worksheet - sheffieldschools** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp (you do not need to worry about this yet) each definition can be written as a set of two conversion factors. they are: 1 mole = molar mass(g) can be written as 1 mole or molar mass (g) **teacher: subject: chemistry-stoichiometry unit grade ...** - interpret coefficients of a balanced equation as mole ratios. use mole ratios from the balanced equation to calculate the quantity of one substance in a reaction given the quantity of another substance in the reaction. (given moles, particles, mass, or volume and ending with moles, particles, mass, or volume of the desired substance)

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