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# Mole Conversion Problem Worksheet Answers

**mole conversion problems - mr. bigler** - mole conversion problems note: some of these are the same compounds as in the "molar mass" worksheet, so you can use the formula weights from that worksheet as the starting point for your calculations. 1. how many moles are in 72.9g of hcl? 2. how many moles are in 79.85g fe 2o 3 3. how many moles are in 11.2' of co 2 gas at s.t.p.? 4. **date name per - chemistry** - date\_\_\_\_\_name\_\_\_\_\_per\_\_\_\_\_ mole conversion problems complete the following practice problems for mole conversion. show your work and units! 1. how many moles are in 72.9 g of hcl? molar mass hcl = 36.46 72.9g 36.46 = 1.999 mol 2. how many moles are in 79.85 g fe 2o3? molar mass = 159.7 79.85g **mole review practice problems - whitney high school** - ii. conversion problems: for each problem you must use dimensional analysis, sig figs, ... 13. which contains more atoms, a mole of copper or a mole of gold? (write down # in each one) 14. which weighs more, a mole of copper or a mole of gold? ... mole review practice problems author: **mole conversion practice - lap - wordpressarku** - simple one step mole conversation problems. the third problem is a two step problem but using the molar masses of elements (which they have done before). the final problem was a two step conversion but for a molecule instead of an element. this last problem is one they have only done once or twice before with help from me. **mole review practice problems - rocklin.k12** - 1 mole/ # grams 1 mole =  $6.02 \times 10^{23}$  atoms & molecules conversion problems: for each problem you must use dimensional analysis & box your answer. 7. given 12 g of aluminum , how many moles do you have? = 2992 26.98 12 g al 41 mol al = 0.44 mol al g al 16. given 34 moles cf 4, how many grams would you have? 34 mol cf 88 g cf 4 =  $3.0 \times 10^3$  g cf 1 ... **mole calculation worksheet - sheffieldschools** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp ( you do not need to worry about this yet ) each definition can be written as a set of two conversion factors. **key - cp - mixed mole problems** - mixed mole problems - key 1. a) how many grams are there in  $1.5 \times 10^{25}$  molecules of co 2? 1.110g 1 mol 44.0 g 6.0210molecules 1.510molecules co1mol 3 23 2 25!=!!!! b) what volume would the co **mole to grams, grams to moles conversions worksheet** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp (you do not need to worry about this yet) each definition can be written as a set of two conversion factors. they are: 1 mole = molar mass(g) can be written as \_\_\_1 mole or \_molar mass (g)

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