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## Mole Conversions Name Chem Worksheet 11 3

**mole conversions worksheet - gorechem.weebly** - mole-particle conversions (use avogadro's number for your conversions) 1. how many moles of magnesium are in  $223.01 \times 10$  atoms of magnesium? 2. how many molecules are there in 4.00 moles of glucose,  $C_6H_{12}O_6$ ? 3. how many moles are  $1.20 \times 10^{25}$  formula units of calcium iodide? 4. how many formula units are in 12.5 moles of calcium phosphate? **mole conversions name chem worksheet 11-3** - the mole is a unit for count, as is the dozen. a dozen is 12 items, but a mole is 602 000 000 000 000 000 000 000, or  $6.02 \times 10^{23}$  particles. using the periodic table we can find the molar mass, or the mass of a mole of a substance. the molar volume is the volume of a mole of substance. all gases have the same molar volume when measured at **name: section: the mole - mr. jones's science class** - name: \_\_\_\_\_ section: \_\_\_\_\_ the mole this lesson is an introduction to the concept of the mole and calculating conversions related to the mole. the best analogy for understanding a mole is the dozen. a dozen is simply a word that describes a numerical value - twelve. **a mole to grams, grams to moles conversions worksheet** - mole to grams, grams to moles conversions worksheet what are the molecular weights of the following compounds? 1) naoh 2)  $H_3PO_4$  3)  $H_2O$  4)  $Mn_2Se_7$  5)  $MgCl_2$  6)  $NH_4_2SO_4$  there are three definitions (equalities) of mole. they are: 1 mole =  $6.02 \times 10^{23}$  particles 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) **mole worksheet 2 - wappingers central school district** - name \_\_\_\_\_ period \_\_\_\_\_ mole worksheet #2 make the following conversions using unit analysis. use a separate piece of paper, show all work, and circle your final answer. (attach this sheet to your work). set a: one step problems: convert to moles: 1.  $12.04 \times 10^{23}$  atoms he 2.  $3.01 \times 10^{23}$  atoms cu 3. **mole calculation worksheet - sheffieldschools** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at stp (you do not need to worry about this yet) each definition can be written as a set of two conversion factors. they are: 1 mole = molar mass(g) can be written as \_\_\_\_\_ 1 mole or \_\_\_\_\_ molar mass (g) **date name per - chemistry** - date \_\_\_\_\_ name \_\_\_\_\_ per \_\_\_\_\_ mole conversion problems complete the following practice problems for mole conversion. show your work and units! 1. how many moles are in 72.9 g of hcl? molar mass hcl = 36.46 72.9g 36.46 = 1.999 mol 2. how many moles are in 79.85 g fe 2o3? molar mass = 159.7 79.85g **mole review practice problems - whitney high school** - 12. how is the mole similar to a dozen? 13. which contains more atoms, a mole of copper or a mole of gold? (write down # in each one) 14. which weighs more, a mole of copper or a mole of gold? (write down how much each one weighs) 15. a. how are the terms "molar mass" and "atomic mass" different from each other? b. how are the terms the ... **worksheet: mole problems name** - chemistry: a study of matter © 2004, gpb 7.6 5. convert 3.00 moles as 2s 3 to grams. 6. how many moles are represented by 11.5 g of  $C_2H_5OH$ ? 7. what is the mass of ...

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